PART A

1 A sample of alcohol has the following four properties:

1) It has a volume of 100 mL.

- 2) It does not conduct electricity.
- 3) Its temperature is 22°C.
- 4) It is flammable.

Which of these are characteristic properties of alcohol?

A)	1 and 2	C)	2 and 4
B)	1 and 3	D)	3 and 4

2 Which one of the following statements about compounds is true?

- A) They are pure substances that form when two or more different atoms are combined.
- B) They are mixtures that can be separated only by physical means.
- C) They always consist of identical atoms.
- D) They always have the characteristics properties of the elements from which they were formed.

- 1) Matter consists of tiny particles that cannot be divided.
- 2) Matter is made up of four elements (earth, water, fire and air).
- 3) Matter consists of separate particles.
- 4) Matter fills the entire space it occupies.

Which of the above statements were made by Democritus?

- A) 1 and 2 C) 2 and 4
- B) 1 and 3 D) 3 and 4
- 4 The atomic mass of the first 18 elements of the periodic table increases as their atomic number increases.

Which one of the following statements best explains this increase in mass?

- A) Atomic mass depends most of all on the number of electrons and protons in each element.
- B) Atomic mass depends most of all on the number of electrons and neutrons in each element.
- C) Atomic mass depends most of all on the number of neutrons in each element.
- D) Atomic mass depends most of all on the number of neutrons and protons in each element.

5 The following are changes undergone by different substances:

- 1) Salt dissolving in water
- 2) Sublimation of iodine
- 3) Oxidation of copper
- 4) Neutralization of an acid by a base

Which of the above are physical changes?

- A) 1 and 2 C) 2 and 3
- B) 1 and 4 D) 3 and 4

6 Two elements X and Y have the following properties.

Element X	Element Y
Metallic lustre	Without metallic lustre
Two valence electrons	Four valence electrons
Twenty neutrons	Six neutrons
Located in the 4th period	Conducts electricity

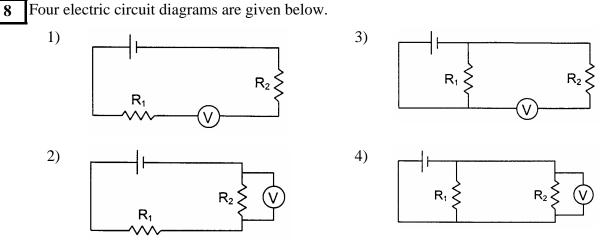
Which symbols from the periodic table correspond to elements X and Y?

A)	K and B	C)	Ca and C
B)	K and C	D)	Ca and B

7 You are given an electromagnet and notice that its magnetic field is not strong enough.

Which one of the following changes will increase the magnetic field strength of this electromagnet?

- A) Replacing the soft iron core with an aluminum core
- B) Increasing the current intensity
- C) Decreasing the number of turns
- D) Increasing the temperature



You wish to measure the potential difference across the terminals of resistor R₂.

C)

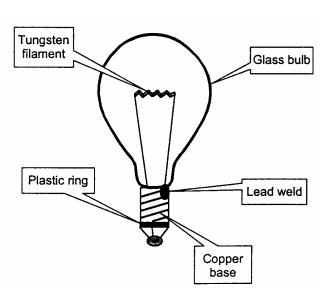
D)

2 and 3

2 and 4

Which diagrams show a correctly connected voltmeter?

- A) 1 and 3
- B) 1 and 4
- 9 The diagram on the right shows the main parts of an incandescent light bulb.



Which of the substances used to make this light bulb are insulators?

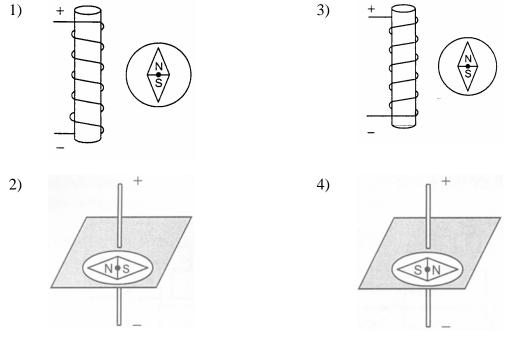
A) Copper and lead

C) Plastic and glass

B) Plastic and lead

D) Glass and tungsten

10 The diagrams below illustrate a compass placed in magnetic field.

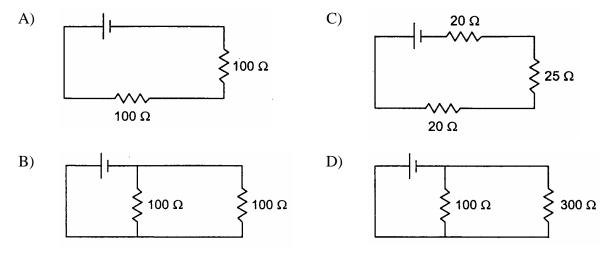


Which diagrams show the compass needle pointing in the correct direction?

A)	1 and 2	C)	2 and 3
B)	1 and 4	D)	3 and 4

11 The four electric circuits given below each consist of a power supply and resistors with different values.

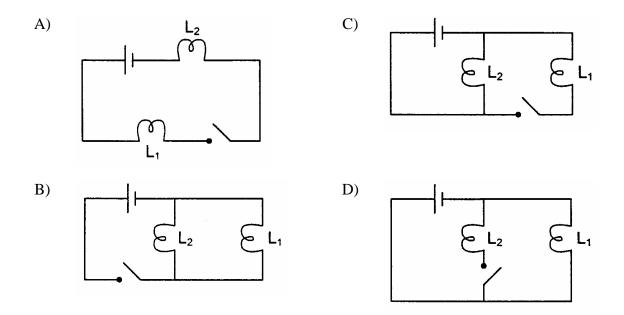
Which circuit has the lowest equivalent resistance?



12 A mystery circuit consists of two light bulbs (L_1 and L_2), a swith, and a power supply. The following table shows what happens to both light bulbs when the switch is opened or closed.

Test	Observations
Open the switch	L ₁ stays on
	L ₂ goes out
Close the switch	L ₁ stays on
	L ₂ comes on

Which diagram correctly represents this mystery circuit?



13 During the summer, the 750-W electric motor of a swimming pool pump is used 24 hours a day for 100 days. Electricity costs \$0.05/kW•h.

How much does it cost to use this motor during the summer?

- A) \$3.75 C) \$90.00
- B) \$37.50 D) \$90 000.00

14 When a 3500-g block of lead was heated, its temperature increased from 20°C to 200°C. The specific heat capacity of lead is $0.13 \text{ J/(g} \cdot ^{\circ}\text{C})$.

How much heat energy was absorbed by this block of lead?

A)	81.9 J	C)	81 900 J
B)	9 100 J	D)	91 000 J

15 To check the electrical conductivity of certain substances, a student used a conductivity apparatus equipped with a light bulb.

Her observations are listed in the following table.

Substances	Observations
HC1	Bright light
CH ₃ OH	No light
MgCl ₂	Faint light
NaOH	Bright light
CH ₃ COOH	Faint light
CCl ₄	No light

Which one of the following groups of substances contains only electrolytes?

A)	CH ₃ OH	C)	CH ₃ OH
	CCl ₄		NaOH
			CH ₃ COOH
B)	HCl	D)	HCl
	MgCl ₂		MgCl ₂
	CCl ₄		NaOH
			CH ₃ COOH

Liquid	pН		
Tap water	6.8		
Lemon juice	2.3		
Human blood	7.3		
Liquid bleach	11		

16 The following table gives the pH value of four liquids.

Which one of these liquids is weakly acidic?

A)	Tap water	C)	Human blood
B)	Lemon juice	D)	Liquid bleach

17 Which one of the following substances is the major contributor to the greenhouse effect?

A) Carbon dioxide (CO_2)	C) N	litrogen dioxide (NO ₂)
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B)Sulfur dioxide (SO2)D)Ozone (O3)

18 In the laboratory, you are asked to prepare 250 mL of a salt solution that will have a concentration of 24% (m/v).

What mass of solute will you need?

A)	6 g	C)	1 042 g
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B) 60 g D) 6 000 g

19 Which one of the following equations is NOT balanced correctly?

- A) $P_4O_{10} + 6 H_2O \rightarrow 4 H_3PO_4$
- B) $2 F_2 + 2 H_2O \rightarrow 4 HF + O_2$
- C) $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$
- D) $HCl + 2 NaHCO_3 \rightarrow 2 NaCl + 2 H_2O + 2 CO_2$

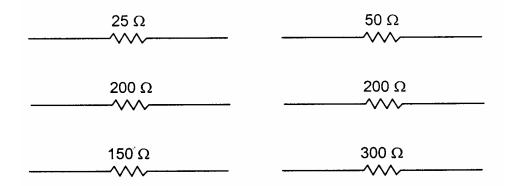
PART B

20 In your answer booklet, draw a simplified atomic model (Bohr-Rutherford) for the potassium atom $^{39}_{19}$ K.

21 In a laboratory, you are given a liquid and asked to determine its density.

- 1. List the materials you would need.
- 2. List the steps in the experiment you would perform in this case (procedure).
- 3. Give an example of how you would calculate the density, using made-up data.

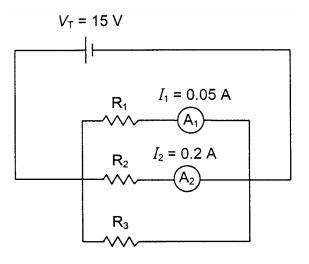
22 In the laboratory, you are given a power supply $(\dashv \vdash)$, conducting wires and the six resistors shown below.



Using the power supply and two of these resistors, you must build a parallel circuit that has an equivalent resistance of 100 Ω .

Draw this parallel circuit in your answer booklet.

23 The following electric circuit consists of a power supply, three resistors (R_1 , R_2 and R_3) and two ammeters A_1 and A_2 .



The potential difference $(V_{\rm T})$ across the power supply is 15 V, ammeter (A) reads 0.05 A and ammeter (A) reads 0.2 A.

The equivalent resistance (R_{eq}) of the circuit is 30 Ω .

What is the value of resistor R₃?

Show all your work.

24 In the laboratory, you are given two acidic solutions. One has a pH value of 5, and the other has a pH value of 6.8. You are also given the following four indicators.

1) Methyl orange

pН	1	2	3	4	5	6	7	8	9	10	11	12	13	14
	Red		O	Orange Yellow										

2) Bromothymol blue

pН	1	2	3	4	5	6	7	8	9	10	11	12	13	14
		Yellow					Green		Blue					

3) Phenolphthalein

pН	1	2	3	4	5	6	7	8	9	10	11	12	13	14
		Colourless							Pink			Dark	pink	

4) m-Cresol purple

pН	1	2	3	4	5	6	7	8	9	10	11	12	13	14
		Yellow						Bro	own			Violet		

Name the only indicator that would allow you to distinguish between the two solutions.

State the colour of that indicator after it is added to each solution.

25 In the laboratory, a student was given a container with 150 mL of a solution of sodium hydroxide (NaOH) that had a concentration of 20 g/L.

She added three 150-mL beakers of water to this solution.

What was the concentration of the diluted solution?

Show all your work.

4. CORRECTION KEY

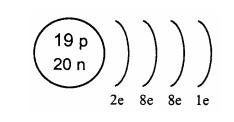
PART A

Questions 1 to 19, 4 marks or 0 marks.

1	С
	A
3	В
4	D
5	A
6	С
7	В
8	D
9	С
10	С
11	В
12	D
13	С
14	С
15	D
16	A
17	Α
18	В
19	D

PART B

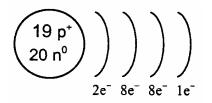
Questions 20 to 25



Example of an appropriate diagram

or

20

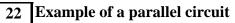


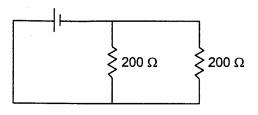
21 Example of an appropriate procedure

- 1. Materials you would need:
 - Balance
 - Graduated cylinder
- 2. Steps in the experiment you would perform (procedure):
 - Weigh the empty graduated cylinder and record its mass.
 - Pour a certain amount of liquid into the graduated cylinder and record the volume of liquid.
 - Weigh the graduated cylinder when it contains the liquid and record its mass.
- 3. Example of how you would calculate the density, using made-up data:

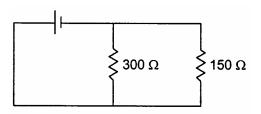
Mass of the cylinder: X Volume of the liquid: Y Mass of the cylinder when it contains the liquid: Z Z - X = mass of the liquid $\frac{\text{mass of the liquid}}{Y} =$ density /4

/4





or





23 Example of an appropriate procedure

Calculation of the total current

$$I_{\rm t} = \frac{V_{\rm T}}{R_{\rm eq}} = \frac{15 \text{ V}}{30 \Omega} = 0.5 \text{ A}$$

Calculation of the current in R₃

$$I_3 = I_t - I_1 - I_2$$

 $I_3 = 0.5 \text{ A} - 0.05 \text{ A} - 0.2 \text{ A}$

$$I_3 = 0.25 \text{ A}$$

Calculation of the voltage in V_3

$$V_{\rm T} = V_1 = V_2 = V_3 = 15 \text{ V}$$

Calculation of the resistance of R₃

$$R_3 = \frac{V_3}{I_3} = \frac{15 \text{ V}}{0.25 \text{ A}} = 60 \Omega$$

Answer

The value of resistor R_3 is 60 Ω .

/4

/4

24 Example of an appropriate procedure

Indicator: 2) Bromothymol blue

Colour of that indicator after it is added to each solution:

The indicator will turn yellow when added to the acid with a pH of 5. The indicator will turn green when added to the acid with a pH of 6.8.

4 marks	Right indicator and correct colours
3 marks	Does not apply
2 marks	Does not apply
1 mark	Right indicator but the colours are incorrect or missing
0 marks	Wrong indicator regardless of the colours

25 Example of an appropriate procedure

Since the volume increases from 150 mL to 600 mL (i.e. it is 4 times greater), the initial concentration will be divided by 4.

$$20 \text{ g/L} \div 4 = 5 \text{ g/L}$$

A procedure involving the formula $C_1V_1 = C_2V_2$ is also acceptable.

Answer

The concentration of the diluted solution was 5 g/L.

/4